



CHAPTER 2: THE STRUCTURE OF THE ATOM



Extra Practice

Section A (Structured Items)

- 1
- (a) Why is the radius of the sodium ion larger than that of the aluminium ion?

[2 marks]
- (b) Why does the sodium atom have a bigger radius than the sodium ion?

[2 marks]
- (c) Which radius is bigger, that of a rubidium ion or that of a sodium ion? Why?

[2 marks]
- (d) There are about 10 million known organic compounds. Explain in terms of bonding, why this occurs.

[2 marks]
- 2
- A calcium atom is electrically neutral but its ion bears a positive charge.
- (a) Explain why.

[2 marks]
- (b) Explain how its ion becomes charged.

[1 mark]
- (c) Why does it become charged?

[2 marks]