



CHAPTER 4: PERIODIC TABLE OF ELEMENTS

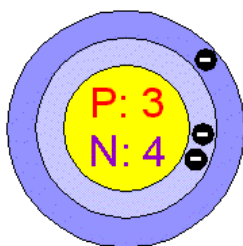


Extra Info

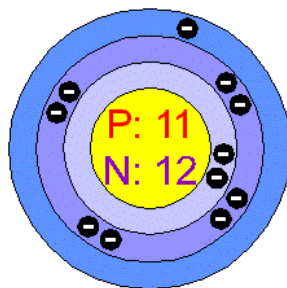
Valence electrons in Group 1

In chemistry, **valence electrons** are the electrons in the outer shell of an atom. Elements with the same number of valence electrons are grouped together in the periodic table of the elements.

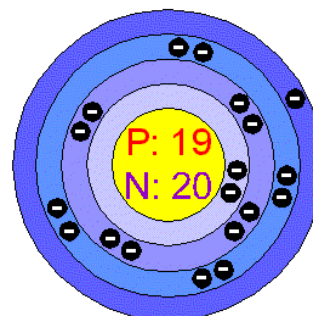
As a general rule, the fewer valence electrons there are in an atom, the more reactive the atom is. The alkali metals, found in group 1 of the periodic table are very reactive metals that do not occur freely in nature. Caesium and francium are the most reactive elements in this group. These metals have only one electron in their outer shells. Therefore, they are ready to lose that one electron in ionic bonding with other elements.



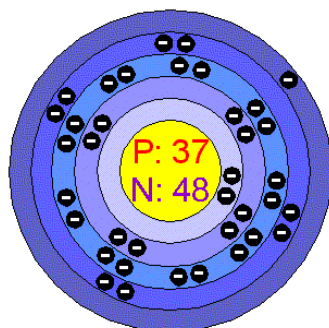
Lithium, Li



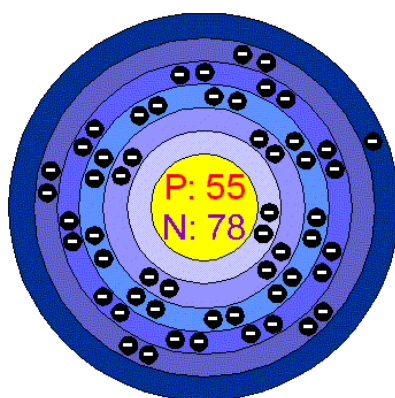
Sodium, Na



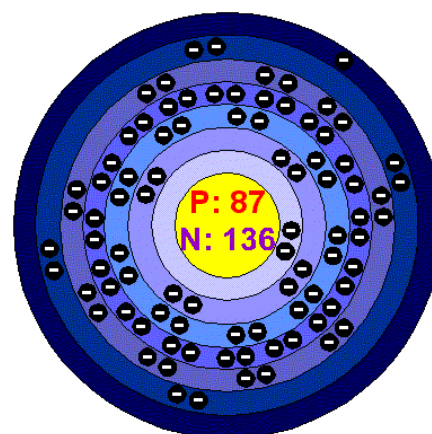
Potassium, K



Rubidium, Rb



Caesium, Cs



Francium, Fr

P = Proton

N = Neutron